

Name: _____ Date: _____ Block: _____

What's in the Beaker? Acids, Bases, Salts, and Buffers

Beaker	Contents
A	30.0 mL of 0.20 M NaOH
B	50.0 mL of 0.30 M HC ₂ H ₃ O ₂
C	50.0 mL of 0.40 M NH ₄ Cl
D	60.0 mL of 0.10 M HCl
E	50.0 mL of 0.50 M NaC ₂ H ₃ O ₂
F	100. mL of 0.20 M NH ₃
G	75.0 mL of 0.20 M NaOH
H	37.5 mL of 0.20 M NaOH
I	90.0 mL of 0.20 M NaOH

Answer Choices
Strong acid
Strong base
Weak acid
Weak base
Acidic Salt
Basic Salt
Neutral Salt
Acidic Buffer
Basic Buffer



QUESTIONS

In the following questions, describe what would be in the beaker (using the answer choices above) when either one of the beakers above is used or a combination of the beakers above is poured together. In addition, calculate the pH of the resulting solution.

#	Question	Answer	pH
1	B		
2	D		
3	B + E		
4	C + F		
5	A + B		
6	B + G		
7	B + I		
8	F + D		
9	A + D		
10	D + G		
11	D + I		
12	A		
13	F		