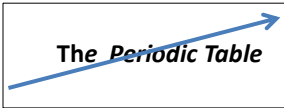


Warm UP 20

small \rightarrow Big

F, N, B, Na

- Organize elements in order of **increasing** atomic radius: N, B, F, & Na
- Ionization Energy is: *E needed to remove/take an e^-*
- Organize elements in order of **increasing** ionization energy: Se, S, Te, O
Te, Se, S, O
- Electronegativity is: *ability to attract e^-*
- Organize elements in order of **increasing** electronegativity: K, Se, Ga, Br
K, Ga, Se, Br
- Summary:



The Periodic Table

Atom Size is: decreasing
 Ionization Energy is: increasing
 Electronegativity is: increasing
- The element that is the **smallest** in size is: He
- The element with the **highest** ionization energy is: He
- The element with the **highest** electronegativity is: F

Luster *Malleable*
Conductor *Ductile*