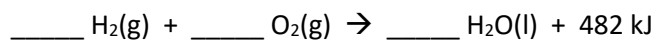


## Warm-Up #45

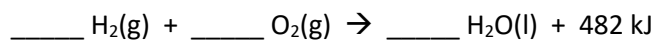
Name: \_\_\_\_\_ Date: \_\_\_\_\_



1. Balance the equation above.
2. Identify the type of reaction.
3. Is the reaction endothermic or exothermic?
4. How many grams of oxygen are needed to produce  $5.32 \times 10^{22}$  molecules of water?

## Warm-Up #45

Name: \_\_\_\_\_ Date: \_\_\_\_\_



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5. What is the theoretical yield, in grams, of water produced when 4.00 moles of hydrogen and 5.00 moles of oxygen are mixed together and react?
  6. What is the limiting reactant?
  7. What is the excess reactant? How many moles of excess remain unreacted?
  8. If in lab, 67.33 grams of water are produced from the reaction mixture in Question #5, then what is the percent yield of water?
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