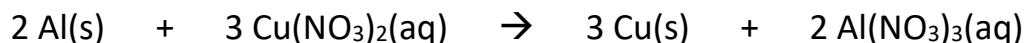


Warm-Up #42

Name: _____ Date: _____

Aluminum reacts with $\text{Cu}(\text{NO}_3)_2$ to produce copper and $\text{Al}(\text{NO}_3)_3$ according to the balanced equation below. 8.0 moles of Al and 9.0 moles of $\text{Cu}(\text{NO}_3)_2$ are placed into a reaction vessel and allowed to react.

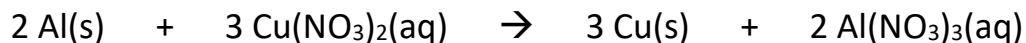
1. Complete a Before-Change-After (BCA) table for the reaction.

**Warm-Up #42**

Name: _____ Date: _____

Aluminum reacts with $\text{Cu}(\text{NO}_3)_2$ to produce copper and $\text{Al}(\text{NO}_3)_3$ according to the balanced equation below. 8.0 moles of Al and 9.0 moles of $\text{Cu}(\text{NO}_3)_2$ are placed into a reaction vessel and allowed to react.

1. Complete a Before-Change-After (BCA) table for the reaction.



2. How many grams of copper are produced?

3. What is the limiting reactant?

4. What is the excess reactant?

5. How many grams of excess reactant remain after the reaction?

2. How many grams of copper are produced?

3. What is the limiting reactant?

4. What is the excess reactant?

5. How many grams of excess reactant remain after the reaction?