

## Warm-Up #36

Name: \_\_\_\_\_ Date: \_\_\_\_\_

I. On the line at the left, write the letter of the type of chemical reaction represented by each equation below.

- \_\_\_\_ 1.  $A + B \rightarrow AB$  a. decomposition reaction
- \_\_\_\_ 2.  $A + BX \rightarrow AX + B$  b. synthesis (combination) reaction
- \_\_\_\_ 3.  $AX + BY \rightarrow AY + BX$  c. double replacement reaction
- \_\_\_\_ 4.  $AB \rightarrow A + B$  d. single replacement reaction

II. Decide whether each of the following equations represents a synthesis (syn), decomposition (decomp), single-replacement (sr), double-replacement (dr), or combustion (combust) reaction. Write your answer on the line and then balance the equation if necessary.

### Rxn Type

- \_\_\_\_\_ 1. \_\_\_\_\_  $BaCl_2 +$  \_\_\_\_\_  $NaOH \rightarrow$  \_\_\_\_\_  $NaCl +$  \_\_\_\_\_  $Ba(OH)_2$
- \_\_\_\_\_ 2. \_\_\_\_\_  $Pb +$  \_\_\_\_\_  $H_3PO_4 \rightarrow$  \_\_\_\_\_  $H_2 +$  \_\_\_\_\_  $Pb_3(PO_4)_2$
- \_\_\_\_\_ 3. \_\_\_\_\_  $CH_4 +$  \_\_\_\_\_  $O_2 \rightarrow$  \_\_\_\_\_  $CO_2 +$  \_\_\_\_\_  $H_2O$

## Warm-Up #36

Name: \_\_\_\_\_ Date: \_\_\_\_\_

I. On the line at the left, write the letter of the type of chemical reaction represented by each equation below.

- \_\_\_\_ 1.  $A + B \rightarrow AB$  a. decomposition reaction
- \_\_\_\_ 2.  $A + BX \rightarrow AX + B$  b. synthesis (combination) reaction
- \_\_\_\_ 3.  $AX + BY \rightarrow AY + BX$  c. double replacement reaction
- \_\_\_\_ 4.  $AB \rightarrow A + B$  d. single replacement reaction

II. Decide whether each of the following equations represents a synthesis (syn), decomposition (decomp), single-replacement (sr), double-replacement (dr), or combustion (combust) reaction. Write your answer on the line and then balance the equation if necessary.

### Rxn Type

- \_\_\_\_\_ 1. \_\_\_\_\_  $BaCl_2 +$  \_\_\_\_\_  $NaOH \rightarrow$  \_\_\_\_\_  $NaCl +$  \_\_\_\_\_  $Ba(OH)_2$
- \_\_\_\_\_ 2. \_\_\_\_\_  $Pb +$  \_\_\_\_\_  $H_3PO_4 \rightarrow$  \_\_\_\_\_  $H_2 +$  \_\_\_\_\_  $Pb_3(PO_4)_2$
- \_\_\_\_\_ 3. \_\_\_\_\_  $CH_4 +$  \_\_\_\_\_  $O_2 \rightarrow$  \_\_\_\_\_  $CO_2 +$  \_\_\_\_\_  $H_2O$

III. Use the solubility table to determine if the each of the following compounds will form a precipitate or aqueous solution.

1. Sodium hydroxide
2. Strontium sulfate
3. Iron (II) chloride
4. Magnesium carbonate

IV. What science class did you sign up to take next year? Do you have any questions about which class is best for you? If you are a senior, then CONGRATULATIONS!

III. Use the solubility table to determine if the each of the following compounds will form a precipitate or aqueous solution.

5. Sodium hydroxide
6. Strontium sulfate
7. Iron (II) chloride
8. Magnesium carbonate

IV. What science class did you sign up to take next year? Do you have any questions about which class is best for you? If you are a senior, then CONGRATULATIONS!