

## Warm-Up #35

Name: \_\_\_\_\_ Date: \_\_\_\_\_

I. On the line at the left, write the letter of the type of chemical reaction represented by each equation below.

- \_\_\_\_ 1.  $A + B \rightarrow AB$  a. decomposition reaction
- \_\_\_\_ 2.  $A + BX \rightarrow AX + B$  b. synthesis (combination) reaction
- \_\_\_\_ 3.  $AX + BY \rightarrow AY + BX$  c. double replacement reaction
- \_\_\_\_ 4.  $AB \rightarrow A + B$  d. single replacement reaction

II. Decide whether each of the following equations represents a synthesis (syn), decomposition (decomp), single-replacement (sr), double-replacement (dr), or combustion (combust) reaction. Write your answer on the line and then balance the equation if necessary.

### Rxn Type

- \_\_\_\_\_ 1. \_\_\_\_\_  $BaCl_2 +$  \_\_\_\_\_  $NaOH \rightarrow$  \_\_\_\_\_  $NaCl +$  \_\_\_\_\_  $Ba(OH)_2$
- \_\_\_\_\_ 2. \_\_\_\_\_  $Pb +$  \_\_\_\_\_  $H_3PO_4 \rightarrow$  \_\_\_\_\_  $H_2 +$  \_\_\_\_\_  $Pb_3(PO_4)_2$
- \_\_\_\_\_ 3. \_\_\_\_\_  $CH_4 +$  \_\_\_\_\_  $O_2 \rightarrow$  \_\_\_\_\_  $CO_2 +$  \_\_\_\_\_  $H_2O$

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III. Use the activity series of metals to decide if a reaction will occur.

1.  $\text{Cu(s)} + \text{Mg(NO}_3)_2(\text{aq}) \rightarrow$  Reaction will occur / No reaction (circle your choice)
2.  $\text{Zn(s)} + \text{CuBr}_2(\text{aq}) \rightarrow$  Reaction will occur / No reaction (circle your choice)
3.  $\text{Mg(s)} + \text{AgNO}_3(\text{aq}) \rightarrow$  Reaction will occur / No reaction (circle your choice)

IV. Use the solubility table to determine if the each of the following compounds will form a precipitate or aqueous solution.

1. Sodium hydroxide
2. Strontium sulfate
3. Iron (II) chloride
4. Magnesium carbonate

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