

Warm-Up #24 Complete the table of ionic compounds.

$$Pb^{4+} Pb^{4+} = 8+$$

| | Name | Cation | Anion | Formula |
|----|-------------------------|-----------|--------------------|-------------------|
| 1. | Calcium fluoride | Ca^{2+} | F^{1-} | CaF_2 |
| 2. | Lead(IV) sulfate | Pb^{4+} | $(SO_4)^{2-}$ | $Pb(SO_4)_2$ |
| 3. | Manganese(III) fluoride | Mn^{3+} | F^{1-} | MnF_3 |
| 4. | Zinc oxide | Zn^{2+} | O^{2-} | ZnO |
| 5. | Magnesium acetate | Mg^{2+} | $(C_2H_3O_2)^{1-}$ | $Mg(C_2H_3O_2)_2$ |

