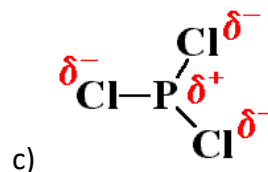
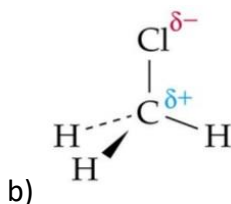
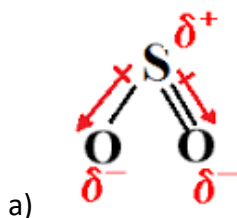


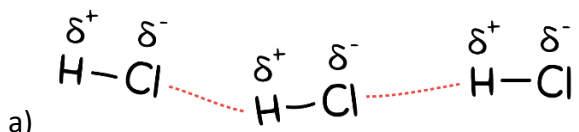
Warm-Up #21

Name: _____ Date: _____

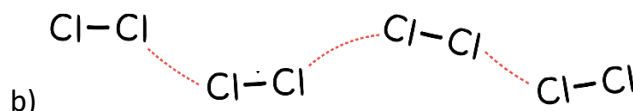
Part 1: Identify each molecule as either polar molecule or nonpolar molecule.



Part 2: For each substance, circle the lines that represent intermolecular forces (IMF's). Name the type(s) of IMF's present in each substance.



Type(s) of IMF's: _____



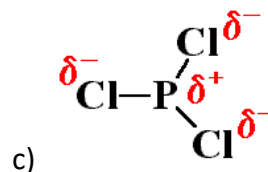
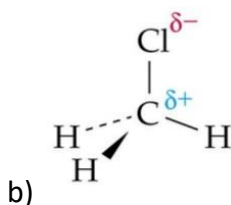
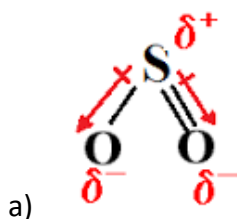
Type(s) of IMF's: _____

Part 3: Which substance from **Part 2** would have the highest boiling point? Explain your reasoning.

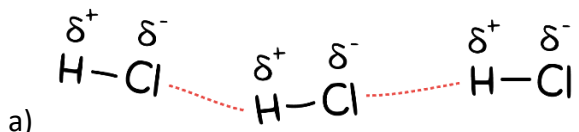
Warm-Up #21

Name: _____ Date: _____

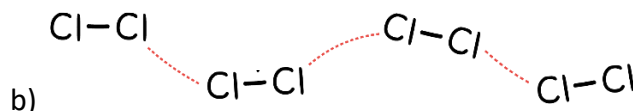
Part 1: Identify each molecule as either polar molecule or nonpolar molecule.



Part 2: For each substance, circle the lines that represent intermolecular forces (IMF's). Name the type(s) of IMF's present in each substance.



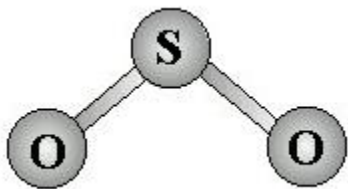
Type(s) of IMF's: _____



Type(s) of IMF's: _____

Part 3: Which substance from **Part 2** would have the highest boiling point? Explain your reasoning.

Part 4: For each molecule below, use electronegativity values to determine bond polarity. Draw δ^+ and δ^- for polar bonds. Determine molecular polarity and type(s) of intermolecular forces.



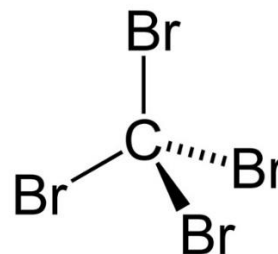
S – O bond polarity: _____

Should you draw δ^+ and δ^- ? _____

Are electrons equally distributed throughout the molecule? _____

Molecular polarity: _____

Type(s) of IMF's: _____



C – Br bond polarity: _____

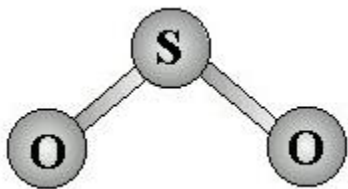
Should you draw δ^+ and δ^- ? _____

Are electrons equally distributed throughout the molecule? _____

Molecular polarity: _____

Type(s) of IMF's: _____

Part 4: For each molecule below, use electronegativity values to determine bond polarity. Draw δ^+ and δ^- for polar bonds. Determine molecular polarity and type(s) of intermolecular forces.



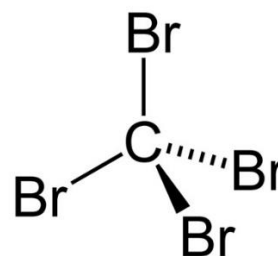
S – O bond polarity: _____

Should you draw δ^+ and δ^- ? _____

Are electrons equally distributed throughout the molecule? _____

Molecular polarity: _____

Type(s) of IMF's: _____



C – Br bond polarity: _____

Should you draw δ^+ and δ^- ? _____

Are electrons equally distributed throughout the molecule? _____

Molecular polarity: _____

Type(s) of IMF's: _____