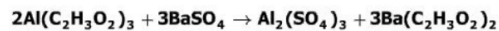


1. Which of these laboratory techniques is best to separate a solid from a liquid to recover the liquid?

- A Titration
- B Chromatography
- C Filtering
- D Vaporization

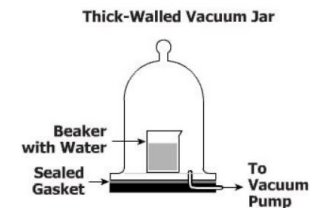
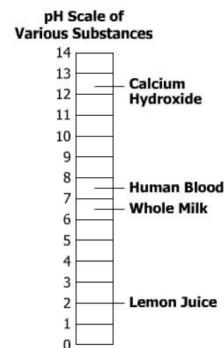


2. Which type of chemical reaction does this equation represent?

- A Synthesis
- B Neutralization
- C Oxidation-reduction
- D Double-replacement

3. Which of the four substances on this pH scale is slightly basic?

- A Calcium hydroxide
- B Human blood
- C Whole milk
- D Lemon juice



4. A beaker of water is placed in a large sealed jar that is attached to a vacuum pump. As air is pumped out of the jar, the water begins to boil because –

- A the temperature of the water decreases as the surrounding pressure decreases
- B the lower pressure inside the jar causes the water to contract
- C the air pressure in the jar has been lowered until it is equal to the vapor pressure of the water
- D the pressure on the water is insufficient to hold the hydrogen and oxygen atoms together, resulting in a decomposition reaction

5. An experiment produced 0.10 g CO₂ with a volume of 0.056 L at STP. If the accepted density of CO₂ at STP is 1.96 g/L, what is the approximate percent error?

- A 110%
- B 92%
- C 71%
- D 8.2%