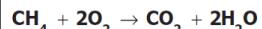


1. If 1.0 mole of methane reacts with oxygen to produce carbon dioxide and water, what mass of water is produced?

A 16 grams
B 18 grams
C 36 grams
D 44 grams

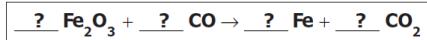


2. The number of molecules in 48.0 grams of oxygen gas is —

F 6.02×10^{23}
G 9.03×10^{23}
H 1.20×10^{24}
J 1.81×10^{24}

3. What are the coefficients of the correctly balanced equation?

A 1, 3, 2, 3
B 0, 2, 2, 3
C 1, 2, 2, 2
D 2, 6, 4, 3



4. A chemist is examining an unidentified element sample with oxidation states of +2, +3, and +6. The element has a shielding effect similar to that of potassium (K). Which statement about the unidentified element is most likely true?
F It has the same number of neutrons as potassium.
G It is a transition metal from the same period as potassium.
H It is one of the heaviest elements in potassium's group.
J It is a mix of three unstable isotopes of potassium.

5. One example of an ionic compound is —

A) F_2
B) CO_2
C) HBr
D) MgCl_2

6. Hydrogen chloride is a covalent compound. Which is a correct Lewis dot structure for HCl? A $\text{:H}:\ddot{\text{Cl}}:$ c $\text{H}::\ddot{\text{Cl}}$

B $\text{H}:\ddot{\text{Cl}}:$ d $\text{H}:\text{Cl}$

7. Le Chatelier's principle describes what happens to a system in equilibrium when a stress occurs. All of the following could shift an equilibrium EXCEPT —

A changing the pressure on the system
B changing the temperature of the system
C changing the identity of the catalyst
D changing the concentration of one of the components

8. A mixture of gases with a pressure of 800 mm Hg contains 10% oxygen and 90% nitrogen by volume. What is the partial pressure of the oxygen gas in the mixture?

F 10 mm Hg
G 80 mm Hg
H 700 mm Hg
J 800 mm Hg

9. The specific heat of aluminum is $0.900 \text{ J/g} \cdot ^\circ\text{C}$. How much heat is required to raise the temperature of a 30.0 g block of aluminum from 25.0°C to 75.0°C .

A 0.540 J
B 1.50 J
C 1350 J
D 1670 J

10. What is the volume of the water in this graduated cylinder?

F 4.39 mL
G 4.41 mL
H 4.55 mL
J 5.61 mL

