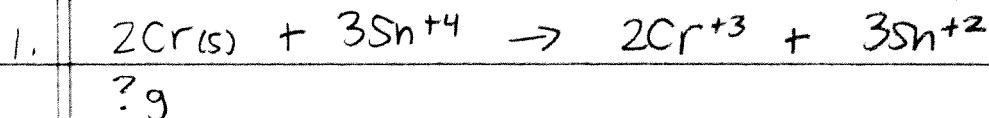
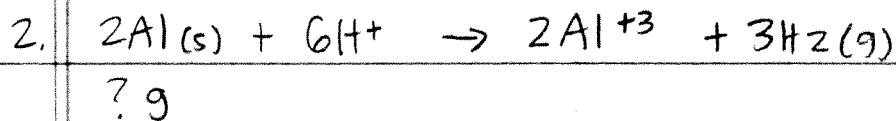


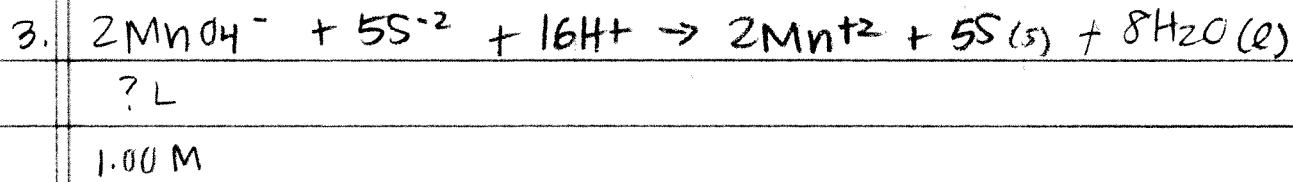
Solution Stoich. w/ REDOX



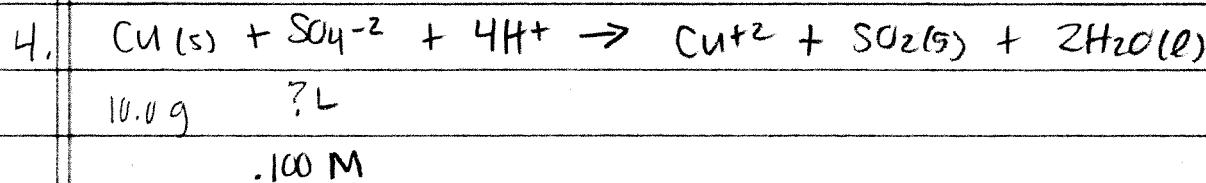
.0200 L $\text{Sn}(\text{NO}_3)_4$	1.00 mol $\text{Sn}(\text{NO}_3)_4$	1 mol Sn^{+4}	2 mol Cr	52.00 g Cr
	1 L $\text{Sn}(\text{NO}_3)_4$	1 mol $\text{Sn}(\text{NO}_3)_4$	3 mol Sn^{+4}	1 mol Cr
= .693 g Cr				



.0400 L HCl	1.00 mol HCl	1 mol H^+	2 mol Al	26.98 g Al	= .360 g Al
	1 L HCl	1 mol HCl	6 mol H^+	1 mol Al	

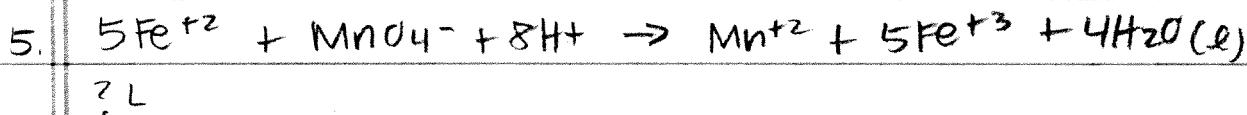


.0300 L Na_2S	2.00 mol Na_2S	1 mol S^{-2}	2 mol MnO_4^-	1 mol KMnO_4	1 L KMnO_4
	1 L Na_2S	1 mol Na_2S	5 mol S^{-2}	1 mol MnO_4^-	1 mol KMnO_4
= .0240 L KMnO_4					

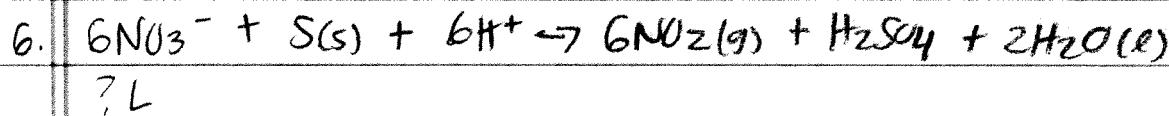


10.0 g Cu	1 mol Cu	1 mol SO_4^{+2}	1 mol K_2SO_4	1 L K_2SO_4	= 1.57 L
	63.55 g	1 mol Cu	1 mol SO_4^{+2}	.100 mol K_2SO_4	K_2SO_4

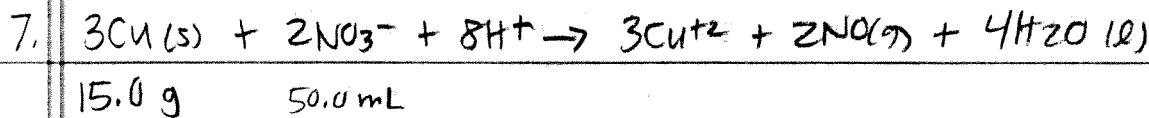
Soln Stoch. w/ REDOX



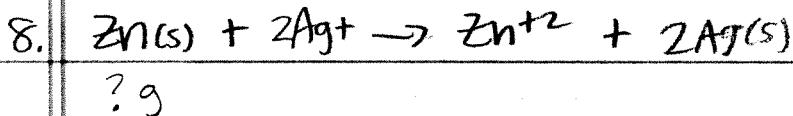
.0500 L KMnO ₄	2.00 mol KMnO ₄	1 mol MnO ₄ ⁻	5 mol Fe ⁺²	1 mol Fe(NO ₃) ₂	1 L
	1 L KMnO ₄	1 mol KMnO ₄	1 mol MnO ₄ ⁻	1 mol Fe ⁺²	.300 M
$= 1.67 \text{ L Fe(NO}_3)_2$					



20.0 g S	1 mol S	6 mol NO ₃ ⁻	1 mol KNO ₃	1 L KNO ₃	$= 12.5 \text{ L KNO}_3$
32.06 g	1 mol S	1 mol NO ₃ ⁻	.300 mol KNO ₃		

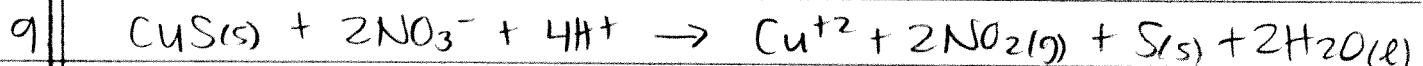


15.0 g Cu	1 mol Cu	2 mol NO ₃ ⁻	1 mol NaNO ₃	$= .157 \text{ mol NaNO}_3 / .0500 \text{ L}$	$= 3.15 \text{ M}$
63.55 g	3 mol Cu	1 mol NO ₃ ⁻			NaNO ₃



.0300 L AgNO ₃	0.400 mol AgNO ₃	1 mol Ag ⁺	1 mol Zn	65.39 g Zn	$= .392 \text{ g Zn}$
	1 L AgNO ₃	1 mol AgNO ₃	2 mol Ag ⁺	1 mol Zn	

Soln Stoich w/ REDOX



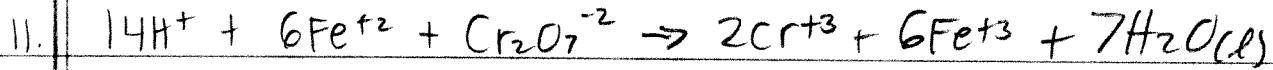
? g 10.0 mL
0.800 M

.0100 L KNO_3	0.800 mol KNO_3	1 mol NO_3^-	1 mol CuS	95.61 g CuS
	1 L KNO_3	1 mol KNO_3	2 mol NO_3^-	1 mol CuS
= <u>.382 g CuS</u>				



30.00 g ? L
1.000 M

30.00 g NO_2	1 mol NO_2	1 mol ClO^-	1 mol NaClO	1 L NaClO
46.01 g NO_2	2 mol NO_2	1 mol ClO^-	1.000 mol NaClO	
= <u>.3260 L NaClO</u>				



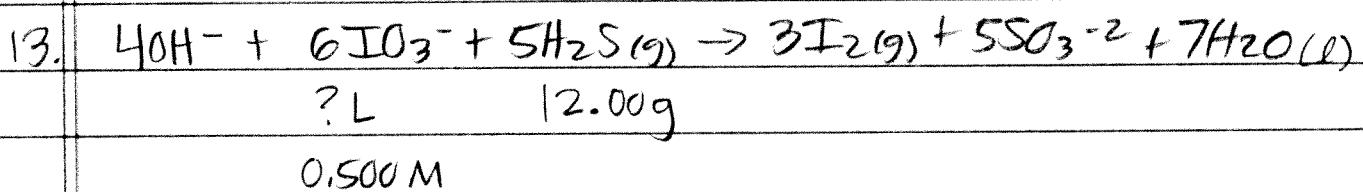
? M 0.200 M
20.0 mL 50.0 mL

.0500 L $\text{K}_2\text{Cr}_2\text{O}_7$	0.200 mol $\text{K}_2\text{Cr}_2\text{O}_7$	1 mol $\text{Cr}_2\text{O}_7^{2-}$	6 mol Fe^{+2}
1 L $\text{K}_2\text{Cr}_2\text{O}_7$	1 mol $\text{K}_2\text{Cr}_2\text{O}_7$	1 mol $\text{Cr}_2\text{O}_7^{2-}$	1 mol $\text{Cr}_2\text{O}_7^{2-}$
= .0600 mol Fe^{+2}			= .0600 mol $\text{Fe}(\text{NO}_3)_2$
	1 mol Fe^{+2}		

$$\frac{\text{M} = \text{mol}}{\text{L}} = \frac{.0600 \text{ mol } \text{Fe}(\text{NO}_3)_2}{.0200 \text{ L } \text{Fe}(\text{NO}_3)_2} = \boxed{3.00 \text{ M } \text{Fe}(\text{NO}_3)_2}$$

Soln Stoch w/ REDOX

12. OMIT, NO EQN.



12.00 g H ₂ S	1 mol H ₂ S	6 mol IO ₃ ⁻	1 mol NaIO ₃	1 L NaIO ₃
	34.08 g H ₂ S	5 mol H ₂ S	1 mol IO ₃ ⁻	0.500 mol NaIO ₃
=	.845 L NaIO ₃			