

Reaction Rate Notes - 3A

Friday, February 15, 2019 9:47 AM

Reaction Rate Notes

- Reaction rate = Speed of rxn.
- In order to react, reactant particles must collide.
- 2 things needed for a "good" collision, in other words to make a reaction occur:
 1. enough energy
 2. proper orientation (Hit in right spot)
- 5 Ways to Increase Reaction Rate (i.e. make the reaction occur FASTER!)
 1. Add more reactants = more collisions = faster rxn
 2. Increase temperature = particles move faster = more collisions = faster rxn
 3. Increase surface area (crush solid into smaller pieces) = more collisions = faster rxn
 4. Change the state of matter: gases react faster than solids.
 5. Add a catalyst = lowers activation energy = faster rxn

