

1. Left of the number, record I if compound is ionic and C if it is covalent.
2. Does it need a Roman numeral??? If yes CIRCLE the question number
3. Complete the name or balanced formula

I 1. $\text{Al}(\text{NO}_3)_3$ Aluminum nitrate

I 2. FeCl_3 Iron (III) chloride

C 3. CS_3 Carbon trisulfide

I 4. TiO_2 Titanium (IV) oxide

I 5. CaCO_3 Calcium Carbonate

I 6. $\text{Cu}(\text{NO}_2)_2$ Copper (II) Nitrite

I 7. $\text{Sn}(\text{CN})_4$ Tin (IV) cyanide

I 8. $\text{Ba}_3(\text{PO}_4)_2$ Barium Phosphate

C 9. N_2O_4 Dinitrogen tetroxide

I 10. Ag_2SO_3 Silver sulfite

I 11. NiSe_2 Nickel (IV) selenide

C 12. P_3O_5 Triphosphorus pentoxide

I 13. $\text{Sn}_3(\text{PO}_4)_2$ Tin (II) phosphate

C 14. N_7Cl_3 Heptanitrogen trichloride

I 15. LiOH Lithium Hydroxide

- I 16. Manganese (III) bromide $\text{Mn}^{3+} \text{Br}_3^{1-}$
- I 17. Calcium acetate $\text{Ca}^{2+} (\text{C}_2\text{H}_3\text{O}_2^{1-})_2$
- C 18. Sulfur dioxide SO_2
- I 19. Tin (IV) sulfate $\text{Sn}^{4+} (\text{SO}_4^{2-})_2$
- I 20. Zinc hydroxide $\text{Zn}^{2+} (\text{OH}^{1-})_2$
- I 21. Lead (IV) nitride $\text{Pb}_3^{4+} \text{N}_4^{3-}$
- I 22. Copper (II) Chlorate $\text{Cu}^{2+} (\text{ClO}_3^{1-})_2$
- C 23. Carbon tetrachloride CCl_4
- I 24. Ammonium phosphate $(\text{NH}_4^{1+})_3\text{PO}_4^{3-}$
- C 25. Diphosphorus pentabromide P_2Br_5
- I 26. Barium cyanide $\text{Ba}^{2+} (\text{CN}^{1-})_2$
- I 27. Cobalt (VI) Chromate $\text{Co}^{6+} (\text{CrO}_4^{2-})_3$
- I 28. Lithium nitrite $\text{Li}^{1+} \text{NO}_2^{1-}$
- C 29. Trisulfur hexafluoride S_3F_6
- I 30. Aluminum Oxide $\text{Al}_2^{3+} \text{O}_3^{2-}$