



#1-4: Write the formula and calculate molar mass.

1. Chlorine
2. Ammonium acetate
3. Aluminum sulfate
4. Carbonic acid



5. Determine the number in the number of moles in 25.00g of chlorine gas. (use the molar mass from # 1)



6. How many particles are in 15 moles of sodium chloride?



7. Determine the number of grams in 3.2×10^{23} particles of carbonic acid (use the molar mass from # 4)



8. How much space in liters will 0.0750 mole of nitrogen gas take up at STP?



9. 60.00 grams of carbon dioxide will take up how much space in liters at STP?