Molar Mass (MM) = mass of one mole of a substance

Example 1)   Molar Mass of $AlCl_3$ = molar mass of one Al + molar mass of three Cl

=   (26.98)   +   (3 x 35.45)

Molar Mass of $AlCl_3$ = **133.33 g/mol**

Example 2)   Molar Mass of $Ba(NO_3)_2$ = MM of one Ba + MM of two N + MM of six oxygens

=   (137.33)   +   (2 x 14.01)   +   (6 x 16)

Molar Mass of $Ba(NO_3)_2$ = **261.35 g/mol**

	Formula	I/M	Name	Molar Mass (g/mol)
1.	$AlCl_3$			(26.98) + (3 x 35.45) = **133.33**
2.			Dinitrogen trioxide	
3.			Lead (II) phosphide	
4.	ClF_3			
5.			Carbon disulfide	
6.	ZnS			
7.	Cu_2SO_4			
8.			Carbon tetrachloride	
9.	$(NH_4)_2CO_3$			
10.			Sulfur hexafluoride	