

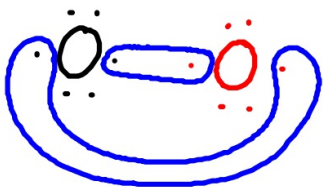
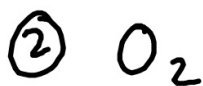
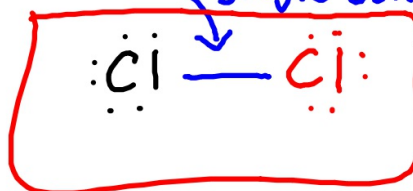
Lewis Structure Notes

- ① Assign correct # of valence e^- to each atom
 - ② Place the atom w/ the fewest # of valence e^- in the center.
 - ③ Connect center atom to every atom w/ a line (one line = single bond = 2 shared e^-)
 - ④ Check octet rule - every atom has 8 e^-
- ★ Exceptions: H needs only 2 e^-
B needs only 6 e^- (Bald head Boron)

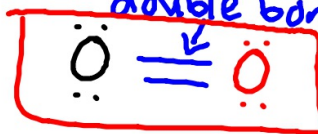
Examples



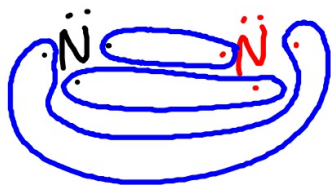
2 shared e^- in covalent bond
single bond



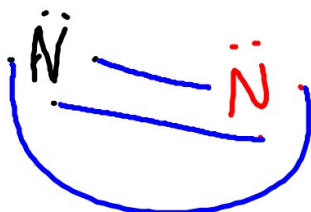
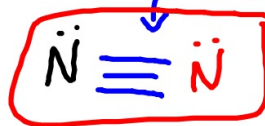
4 shared e^- in covalent
double bond bond



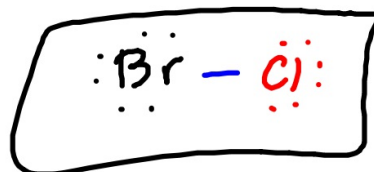
③ N_2



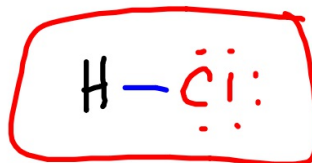
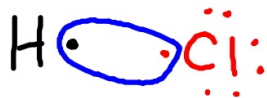
6 shared e⁻ covalent
triple bond bond



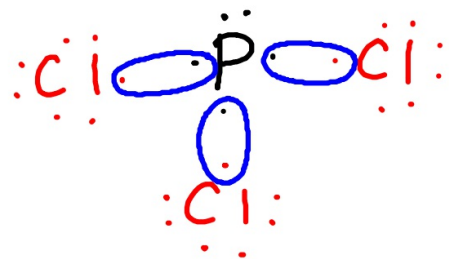
④ $BrCl$



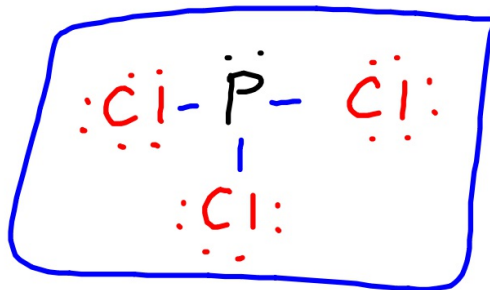
⑤ HCl



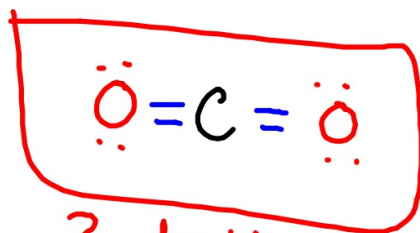
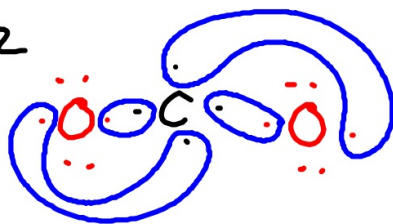
⑥ PCl_3



3 single bonds



⑦ CO_2



2 double bonds