

## Writing Ionic Formula

ionic compound = metal bonded to a nonmetal

= cation bonded to an anion

(+)

(-)

\* Ionic compounds must be neutral (no charge)

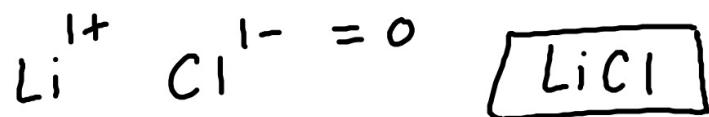
Sum of all positive charge must equal

Sum of all negative charge

\* Do not use prefixes

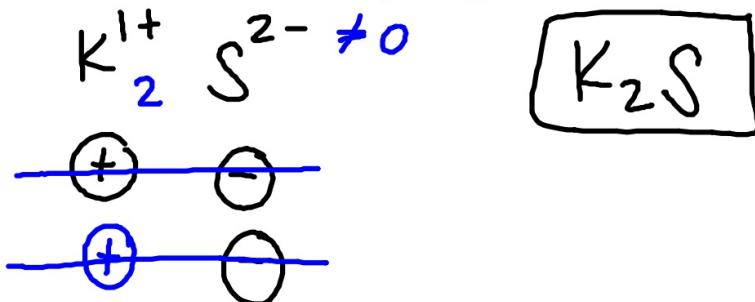
### Examples

a) lithium chloride

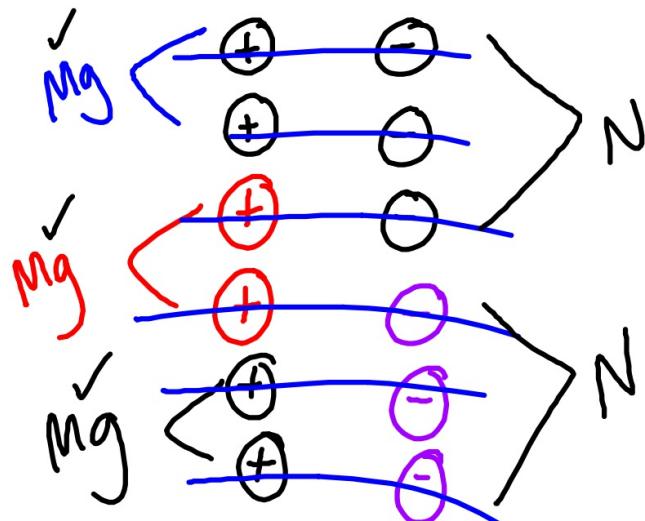
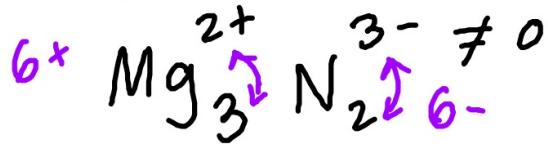


$$(+)(-) = 0$$

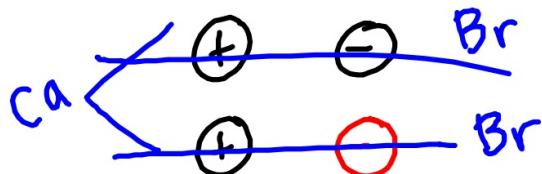
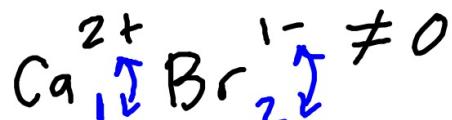
b) potassium sulfide



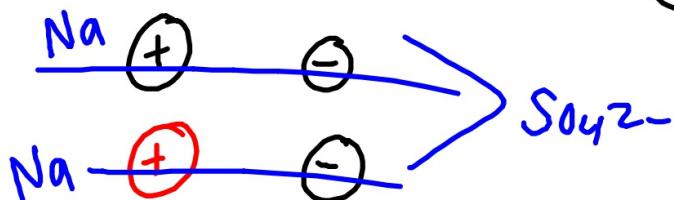
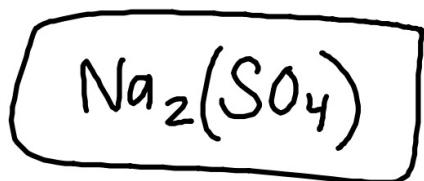
c) magnesium nitride



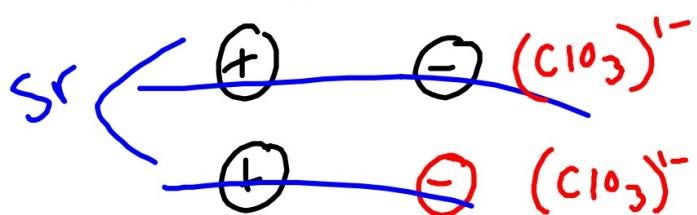
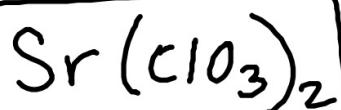
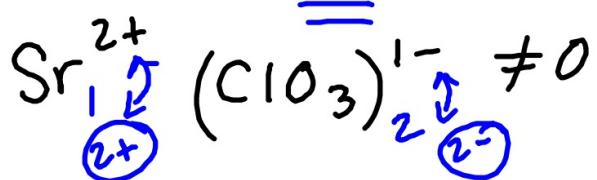
d) calcium bromide

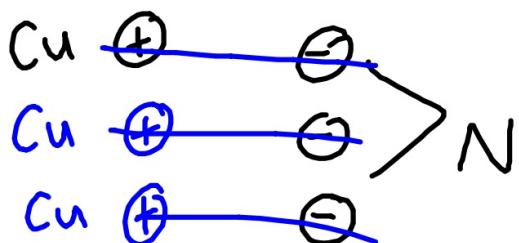
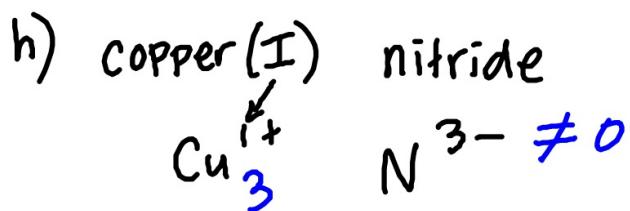
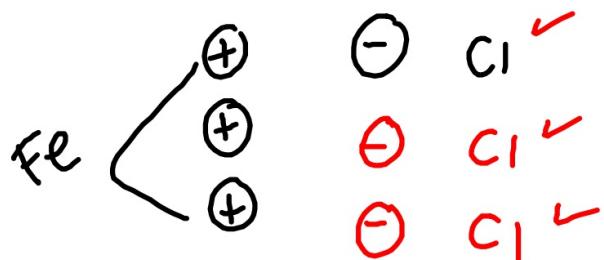
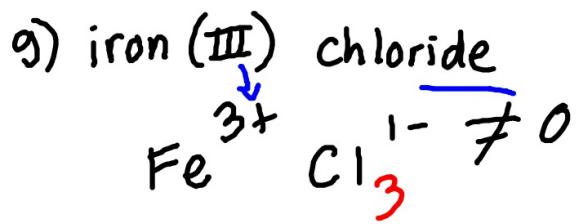


e) Sodium sulfate ← polyatomic ion

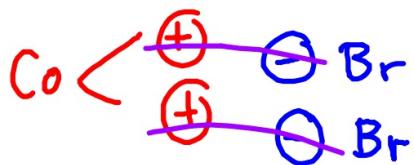
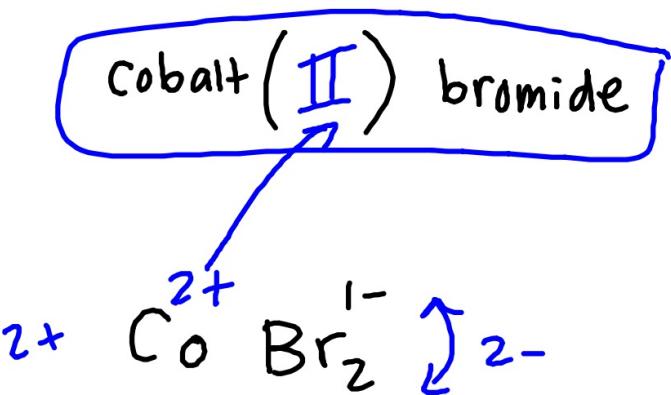


f) strontium chlorate



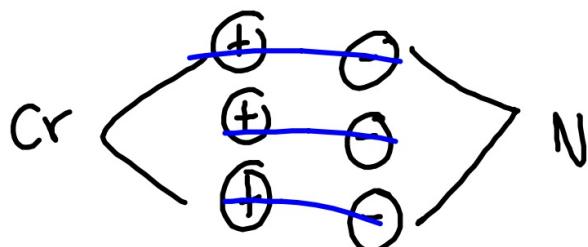
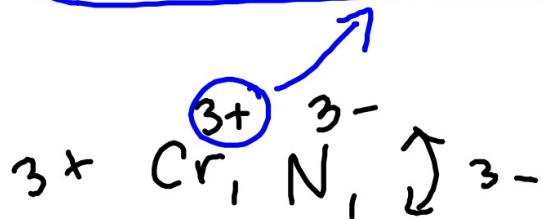


i)  $\text{CoBr}_2$



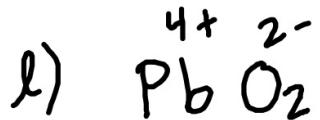
i)  $\text{CrN}$

Chromium(III) nitride





Silver      nitrate



lead(IV) oxide

