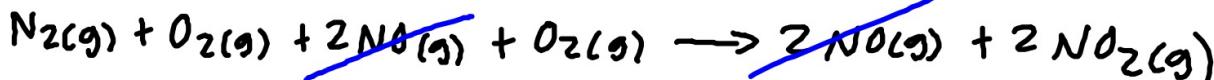
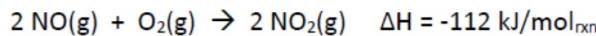
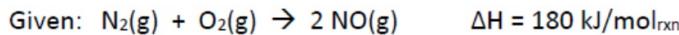
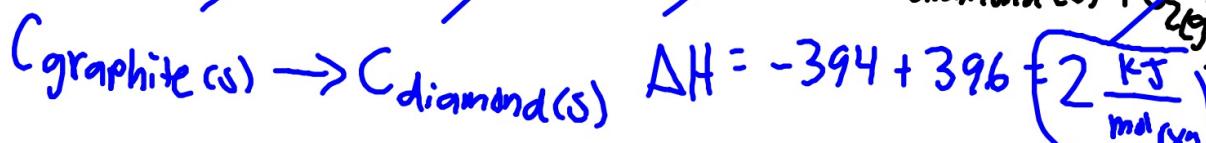
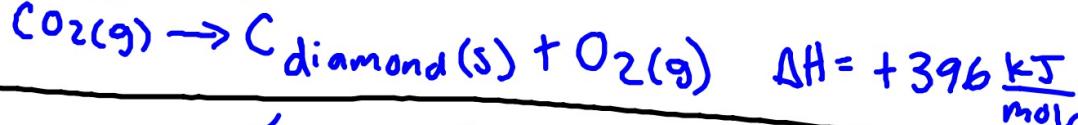
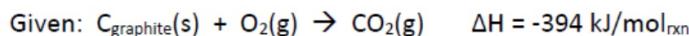


Hess's Law

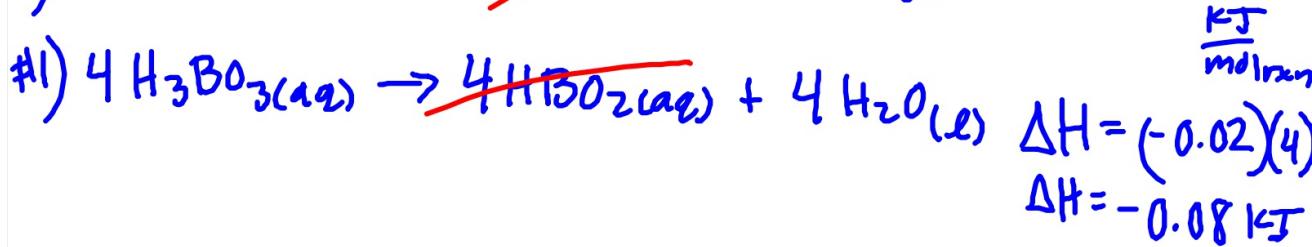
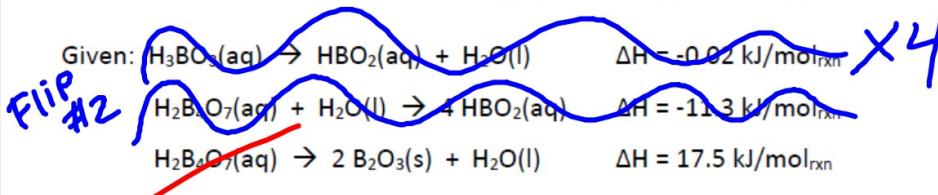
1. Calculate ΔH for the reaction: $N_2(g) + 2 O_2(g) \rightarrow 2 NO_2(g)$.



2. Calculate ΔH for the reaction: $C_{graphite}(s) \rightarrow C_{diamond}(s)$



3. Calculate ΔH for the reaction: $2 H_3BO_3(aq) \rightarrow B_2O_3(s) + 3 H_2O(l)$



$$\Delta H = 17.5 + 11.3 - 0.08 = 28.72 \frac{\text{kJ}}{\text{mol}_{rxn}}$$

Divide By 2

