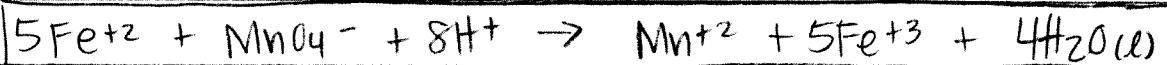
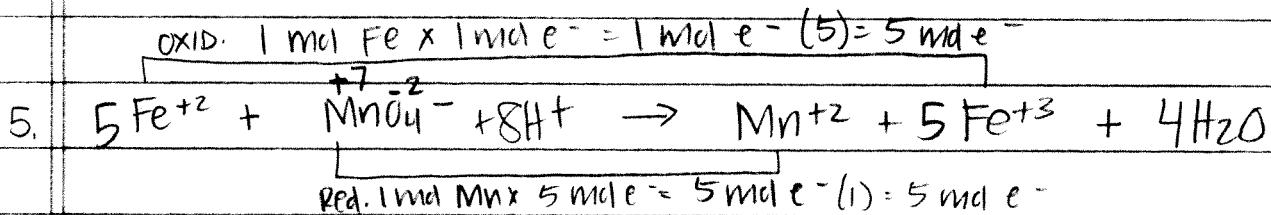
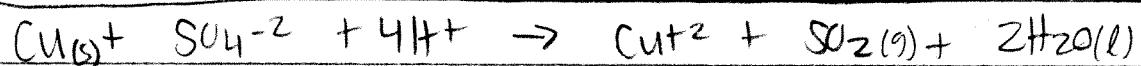
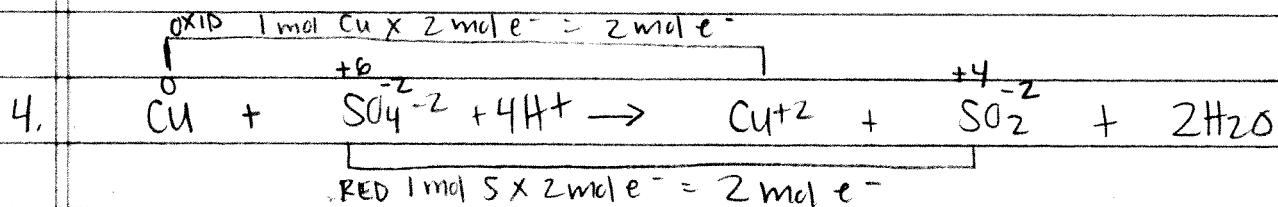
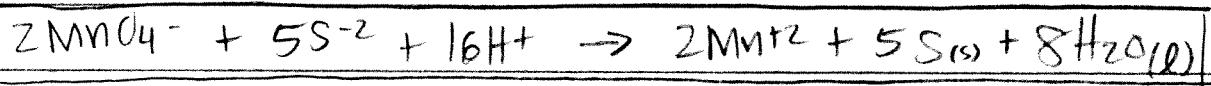
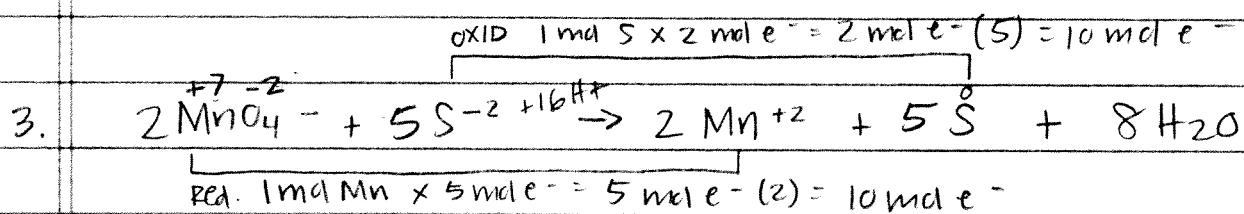
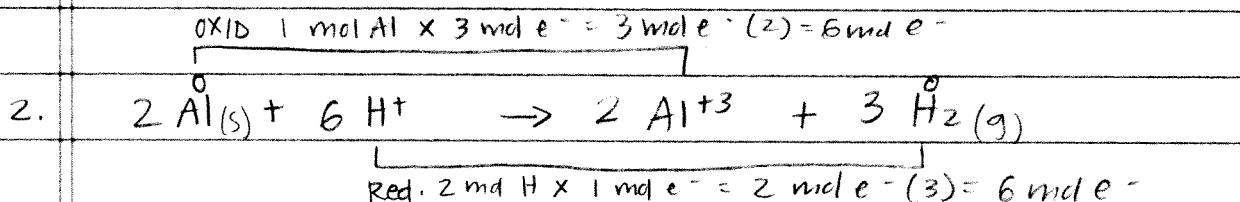
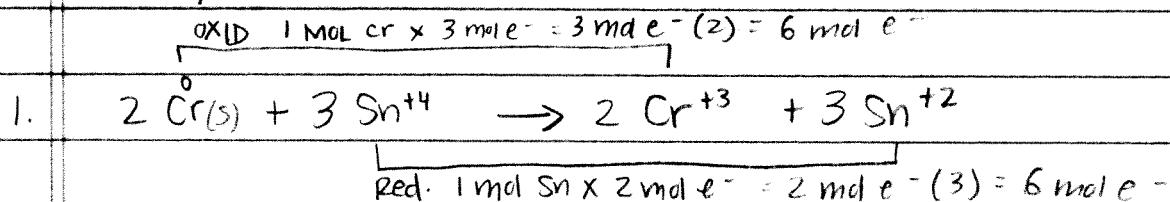
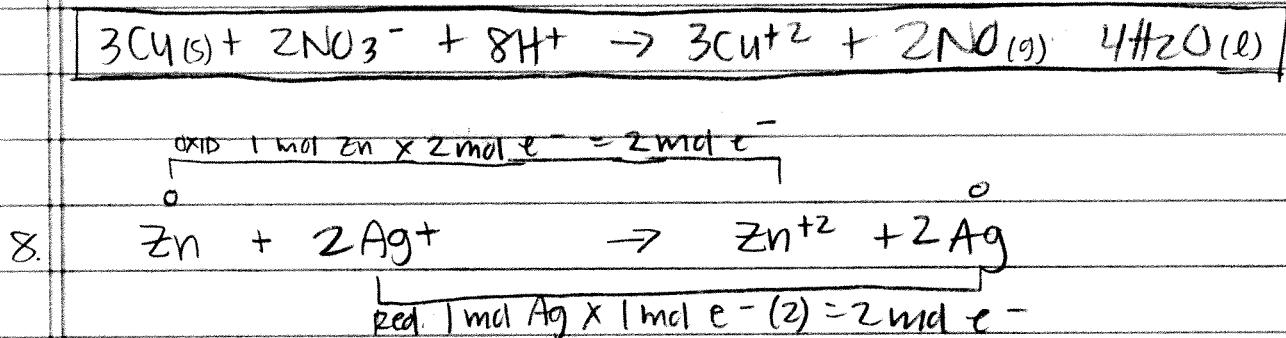
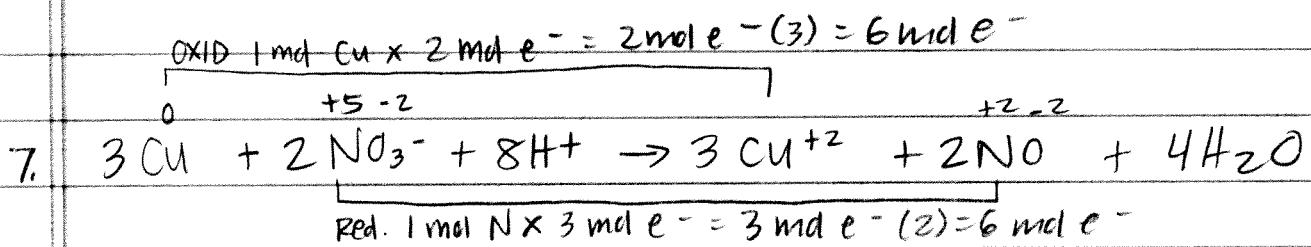
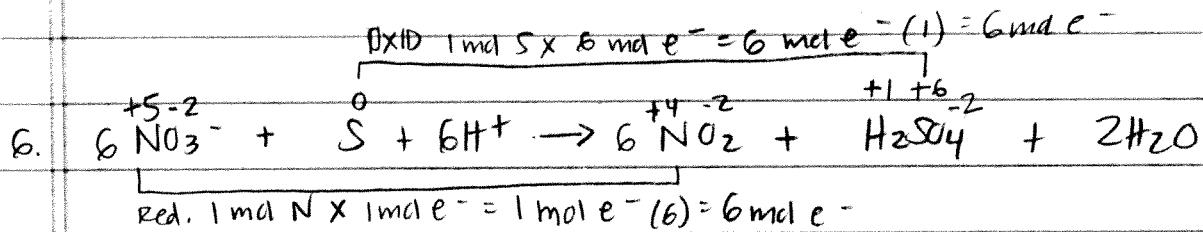


REDOX EQUATIONS

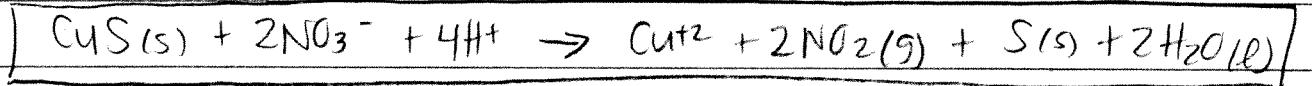


REDOX EQUATIONS



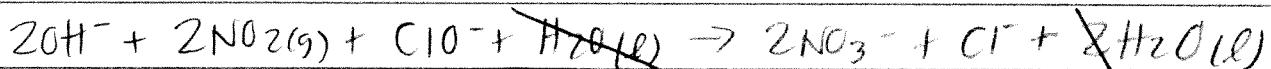
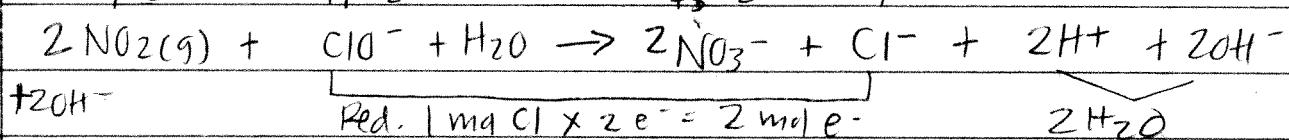
REDOX EQUATIONS

9.



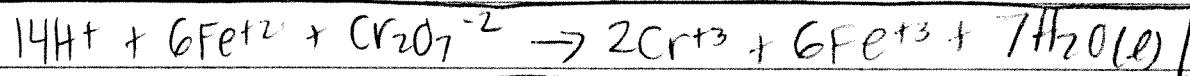
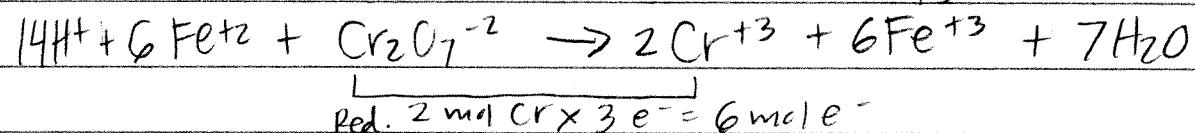
Oxid. $1 \text{ mol N} \times 1 \text{ e}^- = 1 \text{ mole}^- (2)$

3ASIC 10.



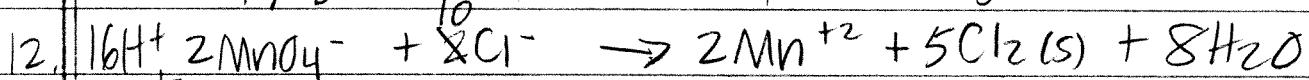
Oxid. $1 \text{ mol Fe} \times 1 \text{ e}^- = 1 \text{ mole}^- (6) = 6 \text{ mole}^-$

11.

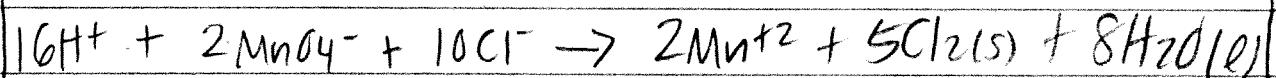


Oxid. $2 \text{ mol Cl} \times 1 \text{ e}^- = 2 \text{ mole}^- (5) = 10 \text{ mole}^-$

12.

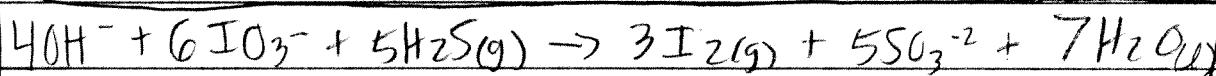
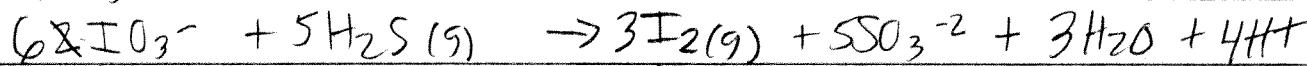


Red. $1 \text{ mol Mn} \times 5 \text{ e}^- = 5 \text{ mole}^- (2) = 10 \text{ mole}^-$

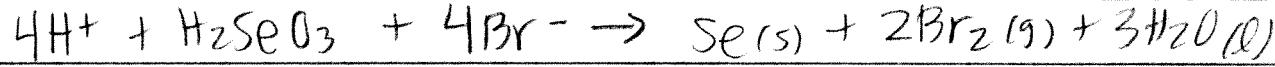
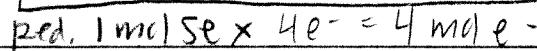
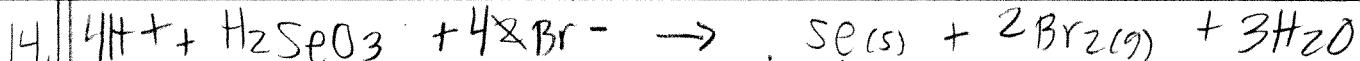


REDOX EQUATIONS

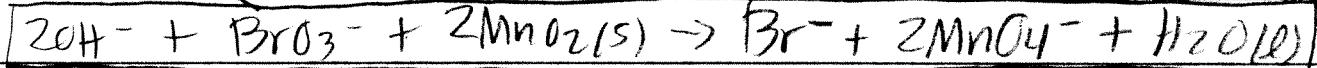
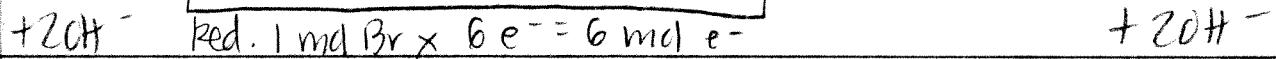
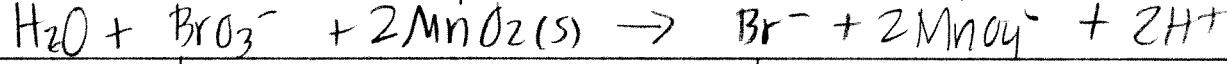
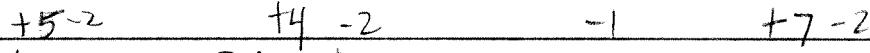
OXID. 1 mol S \times 6e⁻ = 6 mol e⁻ (5) = 30 mol e⁻



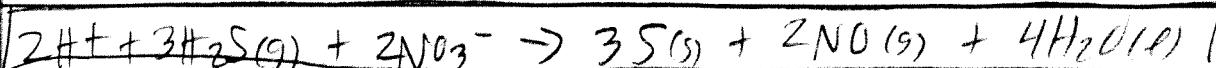
OXID. 2 mol Br \times 1e⁻ = 2 mol e⁻ (2) = 2 mol e⁻



OXID. 1 mol Mn \times 3e⁻ = 3 mol e⁻ (2) = 6 mol e⁻



OXID. 1 mol S \times 2e⁻ = 2 mol e⁻ (3) = 6 mol e⁻



REDOX EQUATIONS

