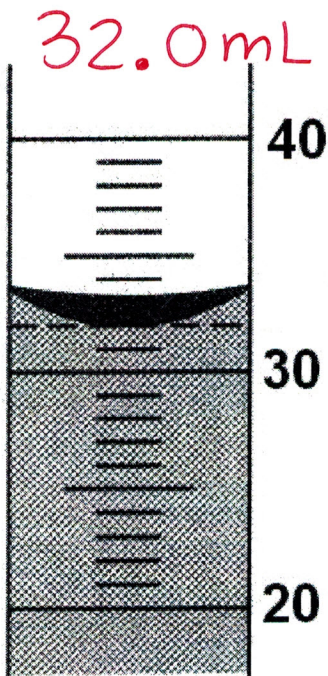
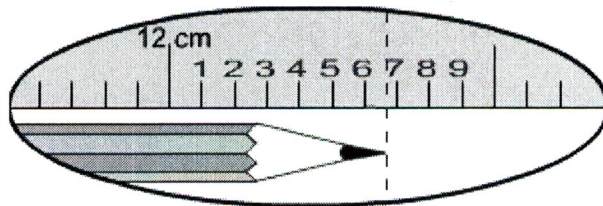


1. Determine the measurement in the following:



12.67 cm



2. Determine the number of sig figs in the following numbers:

a. 401 cg 3

b. 5768 atoms 4

c. 620 L 2

d. 0.0610 kL 3

3. Round each measurement to the specified number of sig figs.

a. 3.8888 cm (3 sig figs) = 3.89 cm

b. 41.27903 mL (2 sig figs) = 41 mL

c. 12035 g (2 sig figs) = 12000 g or  $1.2 \times 10^4$  g

4. Convert the following:

a. 2.75 cm to 27.5 mm  $10^{-2}$   $10^{-3}$  2.75 Right one

b. 562 mm to 0.562 m  $10^{-3}$   $10^0$  562 Left three

c. 0.481 kg to 48100 cg  $10^3$   $10^{-2}$  0.481 Right Five

d. 38.67 mL to 0.00003867 kL  $10^{-3}$   $10^3$  38.67 Left Six  
or  $3.867 \times 10^{-5}$

5. Compare using <, >, or =

a. 1 kL > 0.01L  
1000 L

c. 6.89 cm < 689 m  
0.0689 m

b. 1000 mg = 1g  
1g

d. 0.43 kg > 430 mg  
430000