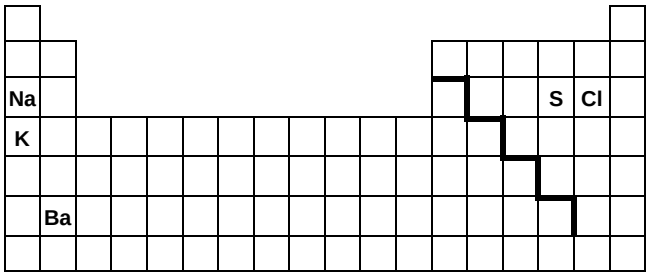
Warm-Up # 26



1. Outline on the periodic table the group of elements that always forms a 2^+ charge.
2. On the periodic table lightly shade the group of elements that always forms a 1^- charge
3. How many electrons does S^{2-} have? _____
4. How many protons does Ba^{2+} have? _____
5. What will be the ion symbol (formula), if the ion has 35 protons and 36 electrons? _____
6. How many electrons does the ion have if it has 12 protons and a 2+ charge? _____
7. Which of the following will gain electrons to form an ion? a) sodium b) argon c) sulfur d) nitrogen
8. If an ion has 15 protons and 18 electrons, what will be its charge? _____
9. Describe with Lewis dot diagrams why potassium and chlorine form an ionic bond.
   (NOTE: Lewis dot diagrams show the valence electrons as dots on the element symbol).

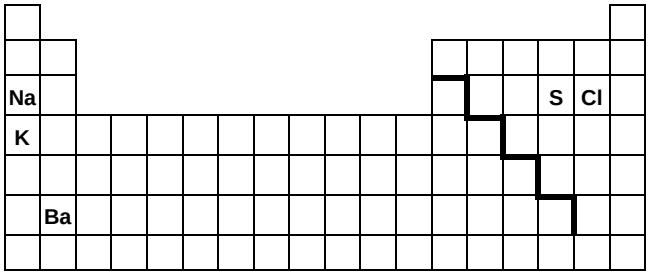
10. Write the names of the following compounds:

a. KF _____

b. Mg_3P_2 _____

c. S_2F_4 _____

Warm-Up # 26



1. Outline on the periodic table the group of elements that always forms a 2^+ charge.
2. On the periodic table lightly shade the group of elements that always forms a 1^- charge
3. How many electrons does S^{2-} have? _____
4. How many protons does Ba^{2+} have? _____
5. What will be the ion symbol (formula), if the ion has 35 protons and 36 electrons? _____
6. How many electrons does the ion have if it has 12 protons and a 2+ charge? _____
7. Which of the following will gain electrons to form an ion? a) sodium b) argon c) sulfur d) nitrogen
8. If an ion has 15 protons and 18 electrons, what will be its charge? _____
9. Describe with Lewis dot diagrams why potassium and chlorine form an ionic bond.
   (NOTE: Lewis dot diagrams show the valence electrons as dots on the element symbol).

10. Write the names of the following compounds:

a. KF _____

b. Mg_3P_2 _____

c. S_2F_4 _____