

Name: _____ Date: _____ Block: _____ **Warm Up #15**

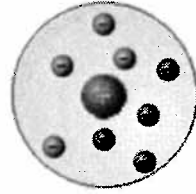
Record the name of the scientist that who responsible for each statement or picture below.

Millikan

1. The charge of the electron is negative and its mass is really small.

Rutherford

2.



Thomson

3. The atom is a positive mass with electrons spread throughout.

Bohr

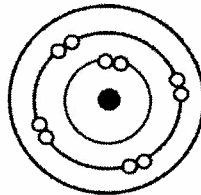
4. Electrons orbit around the nucleus.

Goldstein

5. There is a positive subatomic particle, the proton.

Bohr

6.

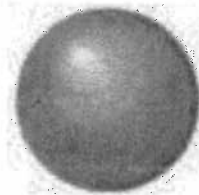


Rutherford

7. The nucleus has a positive dense center.

Dalton

8.

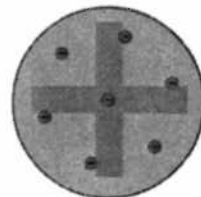


Democritus

9. All matter is composed of really small particles called atoms.

Thomson

10.

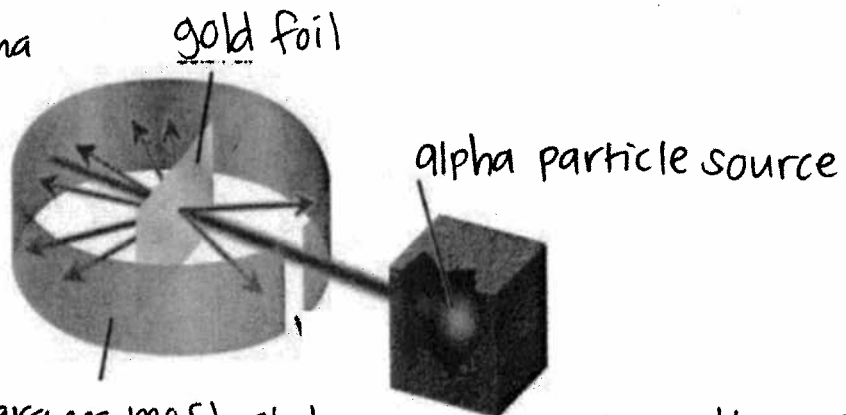


Chadwick

11. Neutron has mass but no charge.

12. Discuss what happened in the gold foil experiment and what conclusions were drawn from the experiment.

Rutherford shot alpha particles (positively charged particles) at a sheet of gold foil. As indicated by the



arrows in the diagrams, most alpha particles went through the gold foil undeflected as expected. However, a very small percentage of alpha particles were deflected. Thus, Rutherford concluded that there must be a very small positively charged nucleus in the atom.

