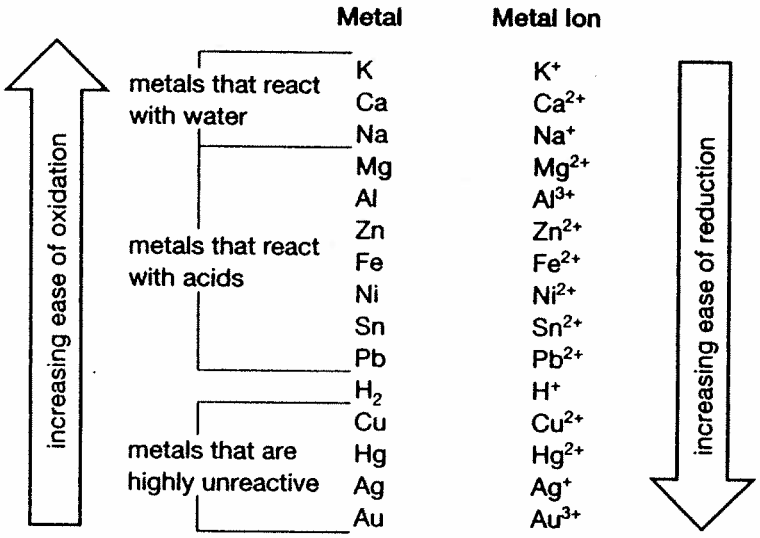
Name _____ Date _____ Class _____

20–2 Review and Reinforcement

Types of Redox Reactions

Activity Series of Metals



Use the activity series above to determine whether a reaction will occur between each of the following reactants. If so, write the balanced equation in the space provided. If not, write "no reaction" on the line provided. If so, circle element that is oxidized, underline the one reduced.

1. $Zn + HgCl_2 \rightarrow$
2. $Ag + Pb(NO_3)_2 \rightarrow$
3. $Cu + AgNO_3 \rightarrow$
4. $Pb + Cu(NO_3)_2 \rightarrow$
5. $Fe + HgCl_2 \rightarrow$
6. $Zn + CaClO_4 \rightarrow$
7. $Ag + Cu(NO_3)_2 \rightarrow$

Answer each of the following questions in the space provided.

8. Describe an activity series. How is it used?

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