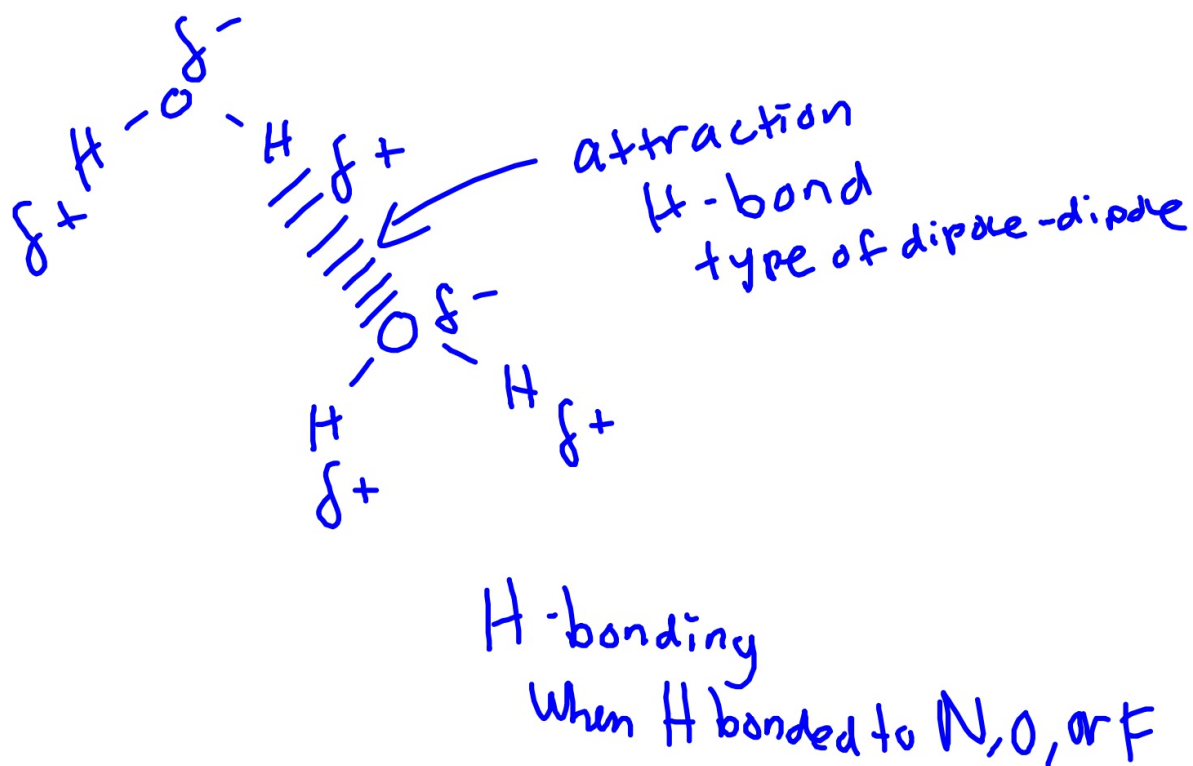
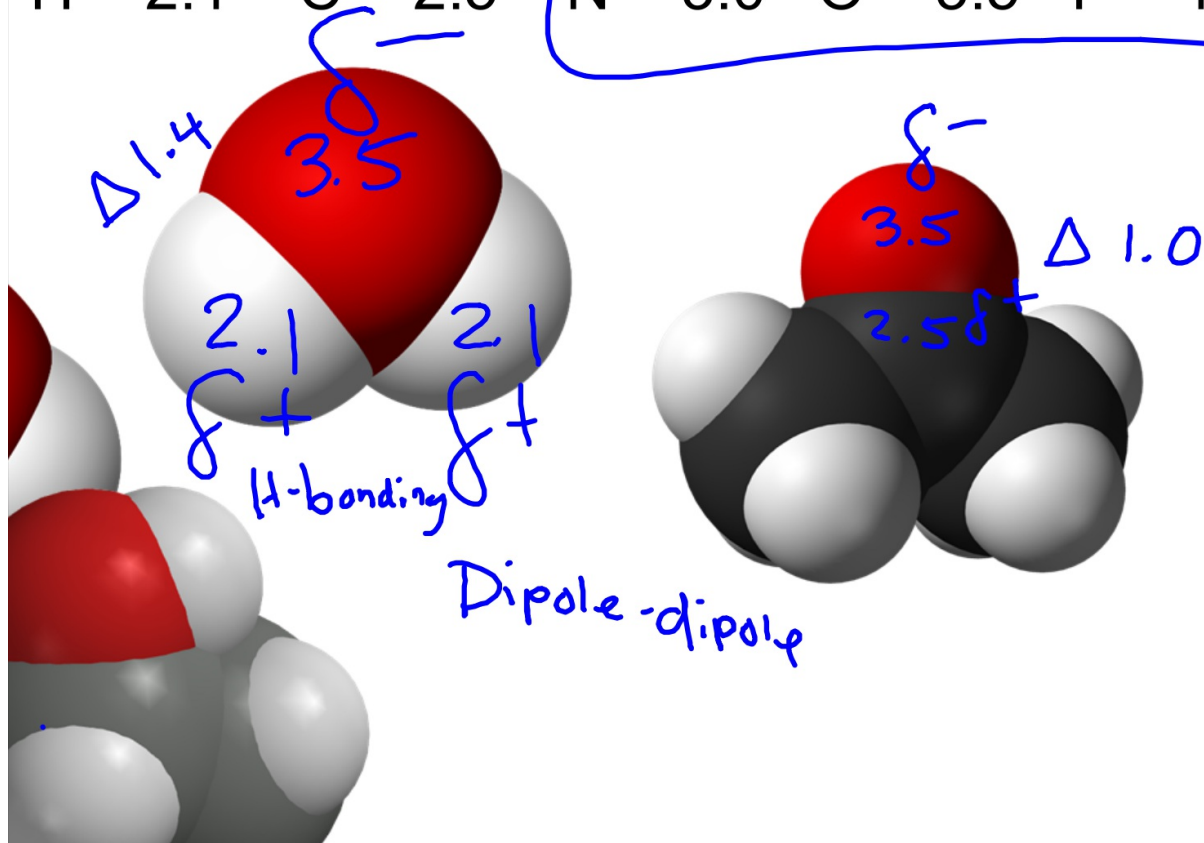
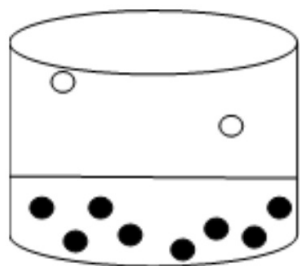


Electronegativity Values

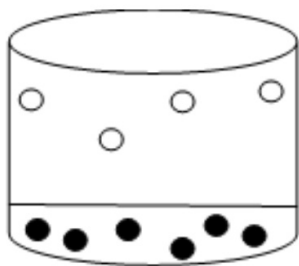
H = 2.1 C = 2.5 N = 3.0 O = 3.5 F = 4.0





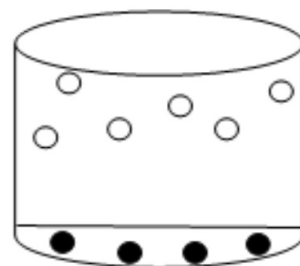
Beaker A

H_2O



Beaker B

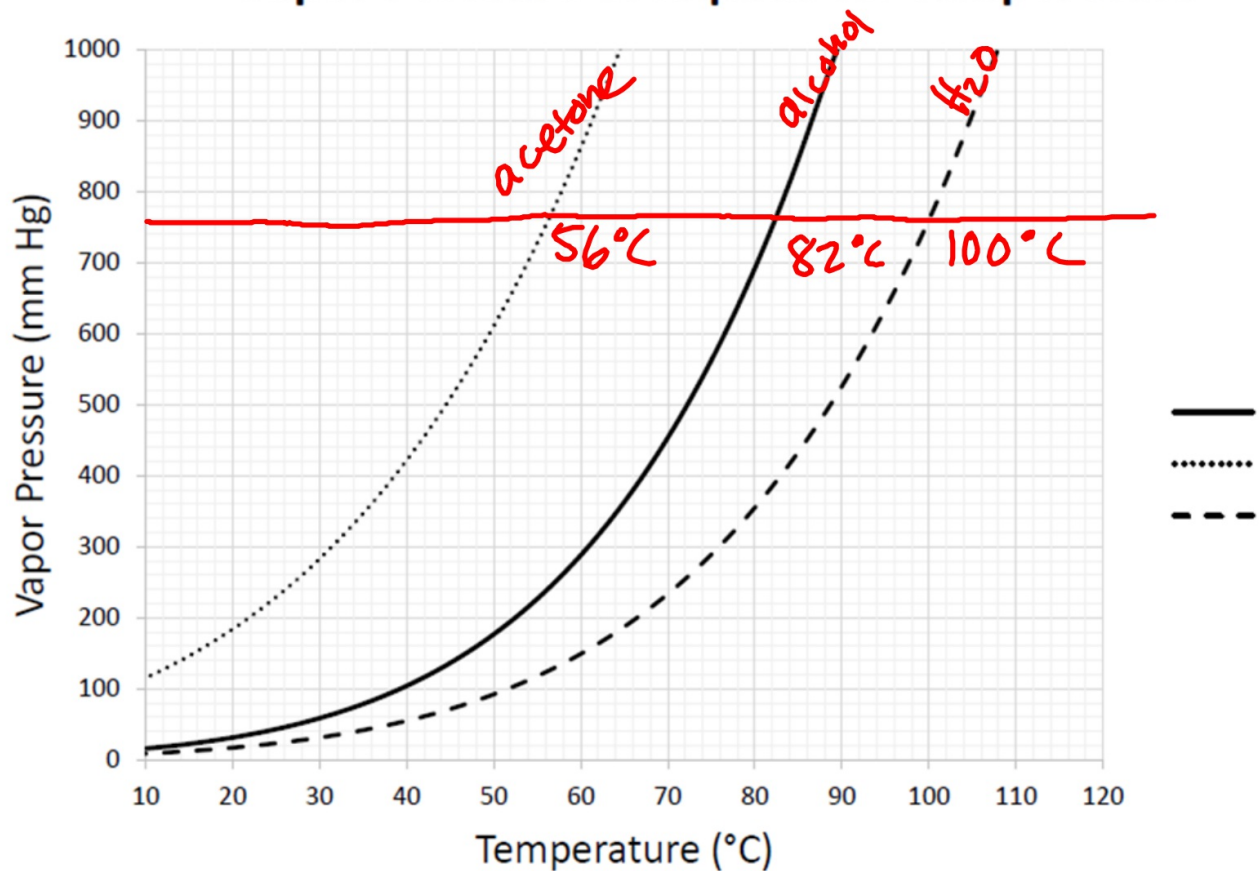
Alcohol
 C_3H_8O



Beaker C

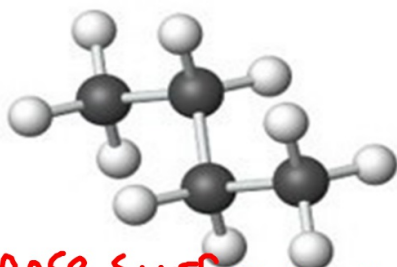
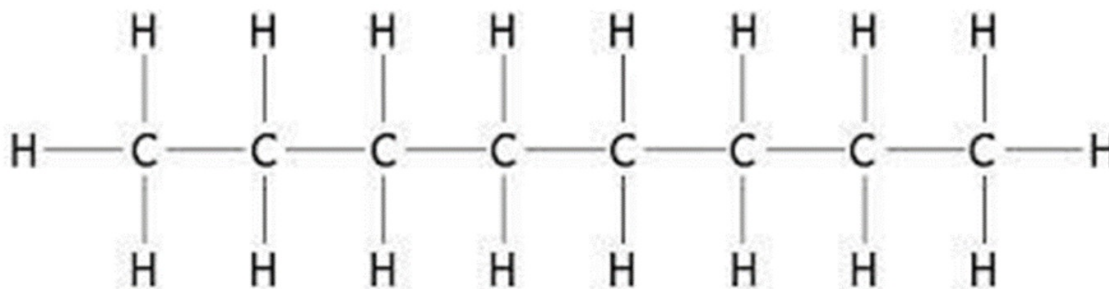
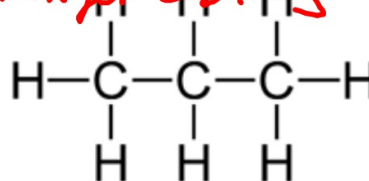
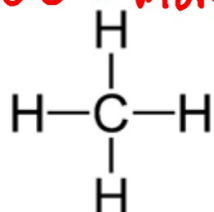
acetone
 C_3H_6O

Vapor Pressure of Liquids vs. Temperature

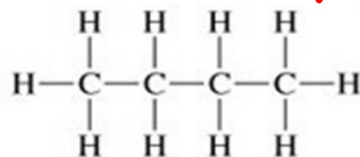


Substance	Molecular Formula	Boiling Point (°C)	State of Matter @ Room Temperature (25°C)
Methane	CH ₄	- 164	Gas
Propane	C ₃ H ₈	- 42	Gas
Octane	C ₈ H ₁₈	125	Liquid

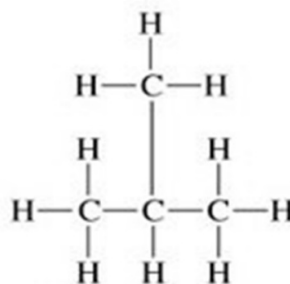
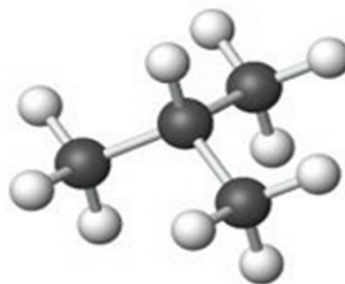
more e⁻ = more polarizable = stronger LDF's



more surface area



Butane, C₄H₁₀



Isobutane, C₄H₁₀

Higher boiling pt
Stronger IMF's

Lower b. p.
Weaker IMF's

