

## D12 HW Net Ionic Equations

Acid/Base, Precipitation, Gas Forming & Redox

- Write the net ionic equation for the following acid/base reactions (*assume all aqueous solutions*):
  - $\text{HCl} + \text{NH}_3$
  - $\text{HF} + \text{LiOH}$
  - $\text{HNO}_3 + \text{Ba}(\text{OH})_2$   
10.0 mL of 1.00M  $\text{HNO}_3$  reacts with 10.0 mL of 1.00M  $\text{Ba}(\text{OH})_2$ . What is the limiting reactant? How many grams of water are made?
- Write the net ionic equation for the following precipitation reactions (*assume all aqueous solutions*):
  - $\text{Pb}(\text{NO}_3)_2 + \text{KI}$
  - $\text{CaCl}_2 + \text{Li}_2\text{SO}_4$   
20.mL of 2.0M  $\text{CaCl}_2$  reacts with 20. mL of 1.0M  $\text{Li}_2\text{SO}_4$ . What is the limiting reactant? How many grams of the precipitate are made?
  - $\text{Sr}(\text{NO}_3)_2 + \text{Na}_2\text{CO}_3$
  - $\text{Zn}(\text{NO}_3)_2 + \text{NaOH}$   
45.0 mL of 1.00M  $\text{Zn}(\text{NO}_3)_2$  reacts with 10.0 mL of 3.00M  $\text{NaOH}$ . What is the limiting reactant? How many grams of the precipitate are made?
- Write the net ionic equation for the following gas forming reactions (*assume all aqueous solutions*):
  - $\text{HBr} + \text{Li}_2\text{CO}_3$
  - $\text{HNO}_3 + \text{K}_2\text{SO}_3$
- Do # 11 on D09 Redox followed by #11 on D10 Redox w/ Soln Stoich.

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- Write the net ionic equation for the following gas forming reactions (*assume all aqueous solutions*):
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